

Stoichiometry and Limiting Reagent:

A reaction equation represent or depict a process in which the reactants or starting materials are converted into products.



→ Symbolizes “yields”

Stoichiometry:

Consider the reaction



This tells the reader that

1 unit of X reacts with 1 unit of Y and yields 1 unit of Z.

Likewise,

1 mole of X + 1 mole Y yields 1 moles Z

Ex: How much Z is produced if 15.0 mol of X is used with excess Y?



From reaction equation

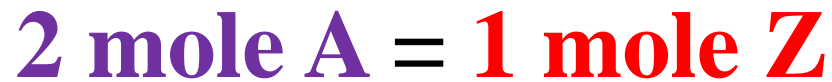
$$\mathbf{1 \text{ mole } X = 1 \text{ mole } Z}$$

$$\mathbf{15.0 \text{ mol } X} \times \frac{\mathbf{1 \text{ mol } Z}}{\mathbf{1 \text{ mol } X}} = \mathbf{15.0 \text{ mol } Z}$$

Ex:2 How much Z is produced if 15.0 mol of A is used with excess B an an alternate synthesis?



From reaction equation



$$15.0 \text{ mol A} \times \frac{1 \text{ mol Z}}{2 \text{ mol A}} = 7.50 \text{ mol Z}$$

Yield:

Actual Yield: The amount of product actually obtained in a synthesis. “Real world result.”

Theoretical Yield: The amount of product that should be obtained based on the amount of reactants used in the reaction. “Expected”

$$\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100 \%$$

Limiting Reagent:

Limiting reagent is the reagent used in the synthesis that is consumed completely and limits how much product is obtained.

Consider the reaction



If 12.5 mol of X is reacted with 11.0 mol of Y.

Consider the reaction



Initial: 12.5 mol X 11.0 mol Y 0 mol Z

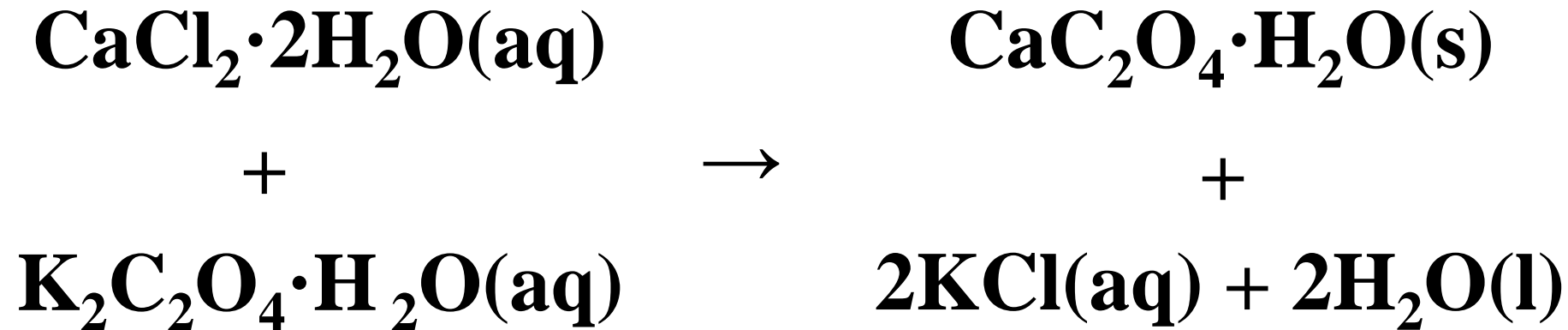
Change: -11.0 mol X -11.0 mol Y +11.0 mol Z

Final: 1.5 mol X 0.0 mol Y 11.0 mol Z

Limiting reagent: Y

Theoretical Yield: 11.0 mol Z

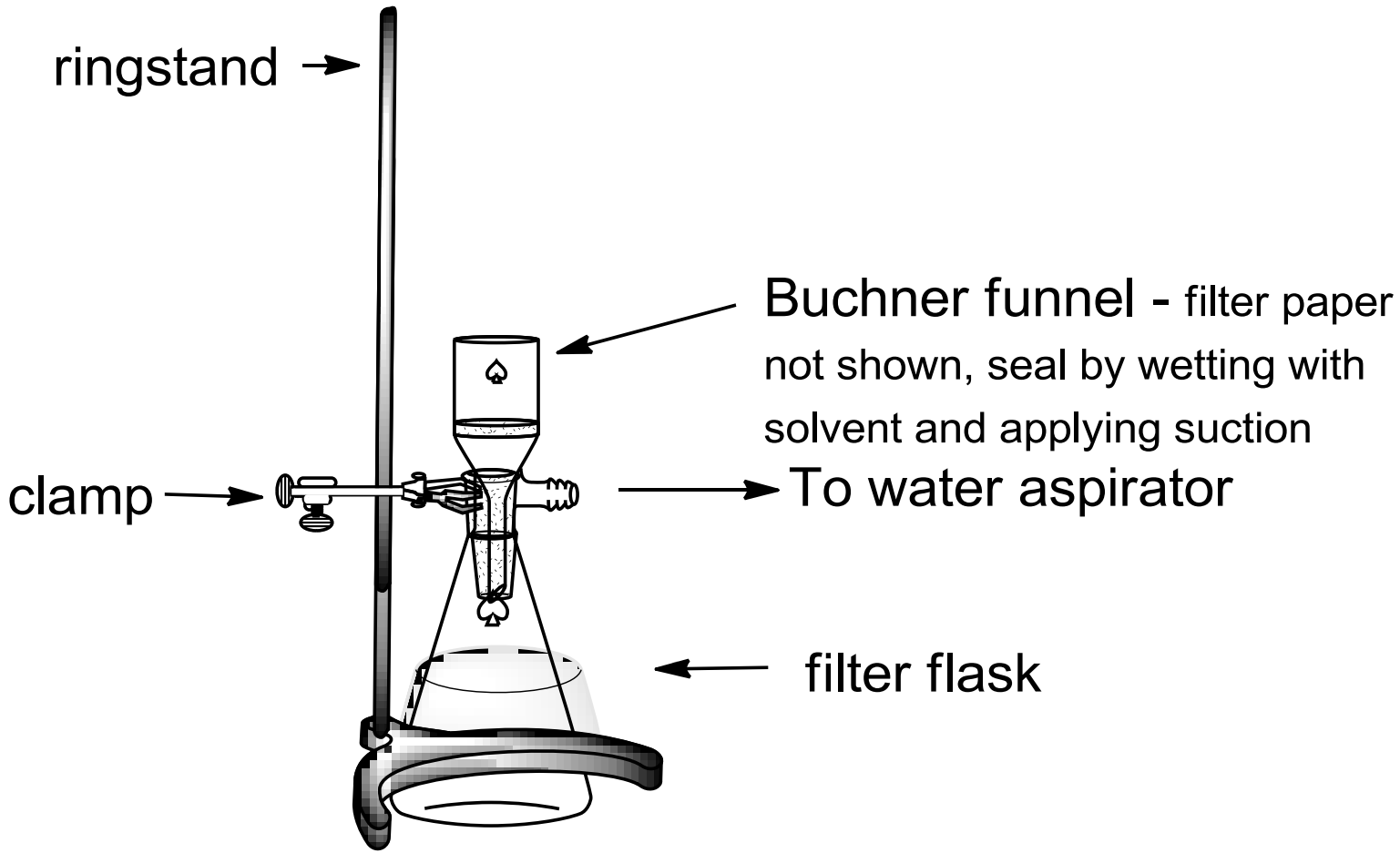
Experiment Reaction:



Will prepare a mixture of the two reactants $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}(\text{s})$ and $\text{K}_2\text{C}_2\text{O}_4 \cdot \text{H}_2\text{O}(\text{s})$ and dissolve in water.

Then predict the limiting reagent and the theoretical yield.

Vacuum Filtration:



Remove rubber tubing connection from aspirator before turning off water